

Video Article

Assembly of a Reflux System for Heated Chemical Reactions

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Abstract

Many chemical experiments require elevated temperatures before any reaction is observed, however heating solutions of reactants can lead to loss of reactants and/or solvent via evaporation if their boiling points are sufficiently low. In order to ensure no loss of reactants or solvent, a reflux system is used in order to condense any vapors produced on heating, and return these condensates to the reaction vessel.

Video Link

The video component of this article can be found at <http://www.jove.com/video/5516/>

Introduction

A reflux system is normally operated using a vertical connection of a water-cooled glass column (reflux condenser) to the outlet of the reaction vessel. This piece of glassware consists of a jacketed column, with water inlet and outlet ports that allow cold tap water to flow through the outer jacket, while reaction vapors are forced through the inner column. The flowing cold water ensures these vapors are condensed on the walls of the inner column and gravity will return these condensates to the reaction vessel. Once a steady reflux has been reached a constant drip of fluid back to the reaction solution should be established. In this fashion, reactions can be left indefinitely without ever needing more solvent to be added. This video will explain the process of connecting your glassware and establishing a steady reflux.

Protocol

1. Preparation of Glassware

1. Ensure the glassware is free of chemical contaminants and if necessary of moisture by allowing to heat in an oven at 100 °C for approximately 30 min.
2. Once cooled, ensure all ground glass joints are contaminant free using either a dry or acetone soaked cleaning tissue specially made for laboratory applications.
3. After dissolving reagents in a suitable solvent in an appropriate reaction vessel (typically a round-bottomed flask) and adding a magnetic stirrer bar, connect the reflux condenser and reaction vessel via the ground glass joint. Attach a Keck clip to the joint.
4. Attach tubing from a cold-water tap to the bottom-most inlet of the reflux condenser and another piece of tubing from the top-most outlet to a sink. Start flowing cold water through the outer jacket of the condenser at such a rate that the sink does not overflow.

2. Heating of Reactants

1. Using a hotplate stirrer, submerge the reaction vessel into a heating bath (oil or water) until the level of the bath is just above the level of the solution within the vessel and hold in place using a ring stand, clamp and boss.
2. Start stirring the reaction. Heat the bath to about 15 °C over the boiling point of the solvent. Once equilibrium between evaporation and condensation has been reached, a steady drip of condensed solvent will start falling back into the reaction vessel from the condenser column.

3. Dismantling the Apparatus

1. Once the desired time has passed, remove the reaction vessel/reflux apparatus from the bath by clamping it higher up the ring stand.
2. Allow the water to continue flowing through the reflux condenser until the solution has reached room temperature.
3. Turn off the flow of water and disconnect the condenser from the reaction vessel.
4. Empty any remaining water in the condenser into the sink and remove tubing from inlet and outlet.

Representative Results

The outcome can be observed after spectroscopic characterization of the resultant solution, as the two reagents should now have reacted to form a new product. Typically, various purification strategies will be required to separate the desired product from undesired side reactions.

In this example, a transesterification reaction between dimethylterephthalate (DMTP) and ethylene glycol has occurred to afford bis-(2-hydroxyethyl)terephthalate and methanol (**Scheme 1**). The refluxing solvent will be the methanol that is being produced (b.p. 65 °C). After heating the starting material (**Figure 1**) under reflux for 45 minutes, nuclear magnetic resonance (NMR) spectroscopy can be used to ensure product formation, as shown in **Figure 2**.



Scheme 1. Transesterification reaction between dimethylterephthalate and ethylene glycol.

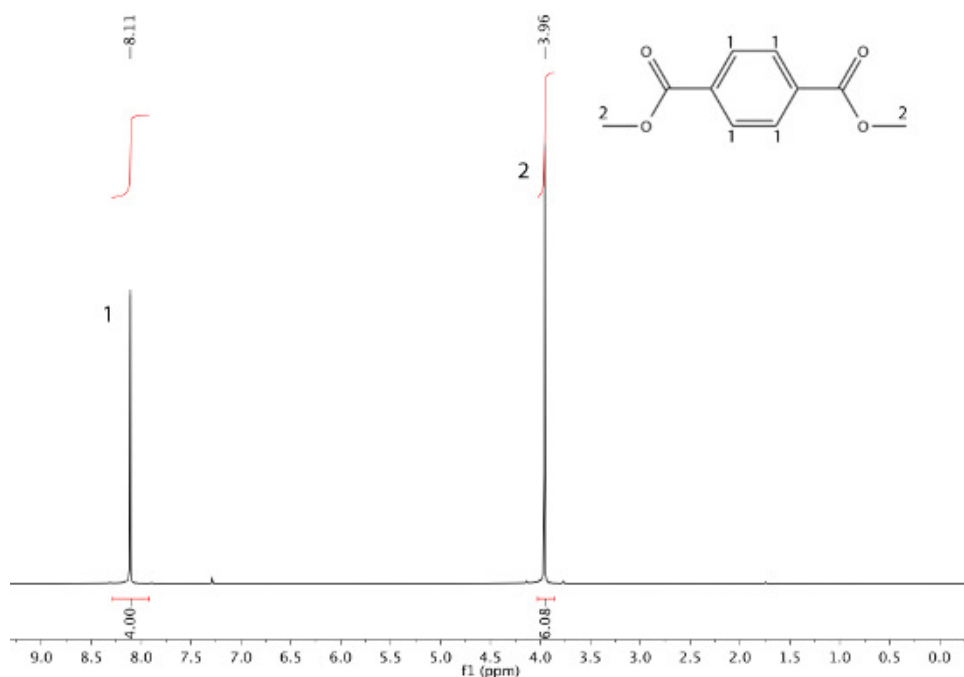


Figure 1. ^1H NMR spectrum of starting material: dimethylterephthalate (DMTP).

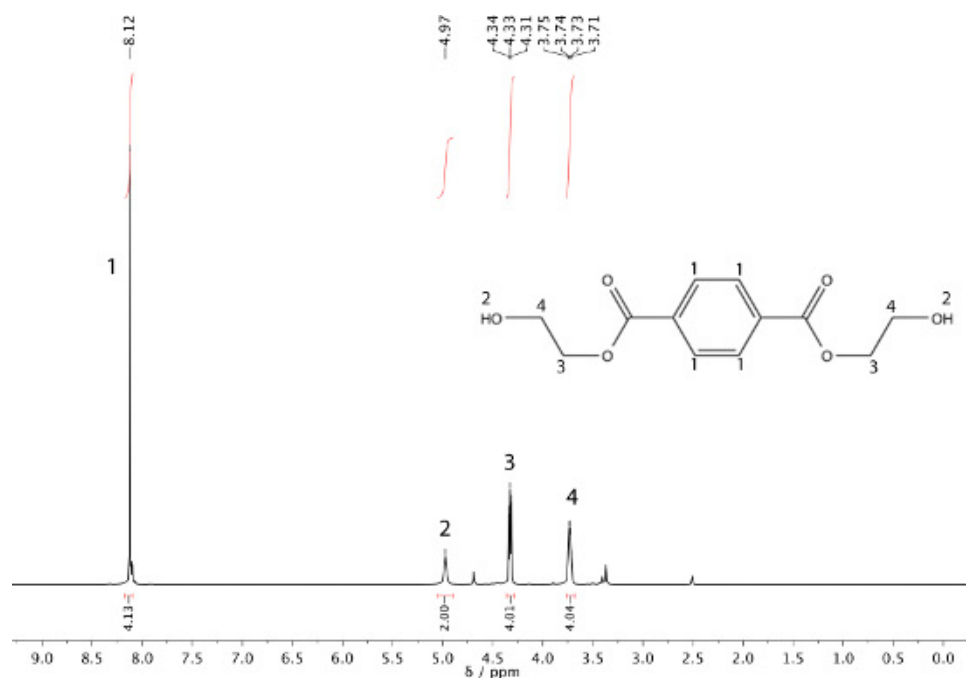


Figure 2. ^1H NMR spectrum of product: bis-(2-hydroxyethyl)terephthalate.

Discussion

Performing reactions under reflux is an important technique to understand. Apart from providing a system whereby solvent and volatile reagents are recycled, it also allows fine control of reaction temperature, as this will be held constant at the boiling point of the chosen solvent. By careful choice of solvent, one can control the temperature within a very narrow range.

More advanced techniques can utilize refluxing solvents to perform sophisticated purification techniques such as Soxhlet extractions or fractional distillation. The later of which is used industrially on a huge scale, for instance in petroleum refineries in order to separate crude oil into various petrol fractions of different boiling points.

Disclosures

No conflicts of interest declared.

References

1. Krell, E. *Handbook of Laboratory Distillation*. Berlin: Elsevier (1982).