

**JoVE: Science Education**  
**n-Butyllithium Titration**  
--Manuscript Draft--

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**Science Education Title:** *n*-Butyllithium Titration [and Addition of \*n\*-Butyllithium to Benzaldehyde](#)

### Overview:

This experiment will demonstrate a simple technique to titrate and obtain an accurate concentration of the organolithium reagent, *n*-butyllithium (*n*-BuLi). Organolithium reagents are extremely air- and moisture-sensitive and proper care must be taken to maintain the quality of the reagent so that it may be used successfully in a reaction. The *n*-BuLi titration experiments should be performed regularly to obtain accurate concentrations prior to usage in a chemical reaction. [Subsequently, we will demonstrate the addition of the titrated \*n\*-BuLi to benzaldehyde.](#)

### Principles:

Organolithium reagents are compounds containing a C–Li bond and are very strong bases. [Because organolithiums are extremely air- and moisture-sensitive, proper care must be taken to handle these reagents.](#) They are sensitive to acids, corrosive, and often pyrophoric. *n*-BuLi, an alkyl lithium, is typically stored at low temperatures (less than 0 °C) [and stored as a solution in hydrocarbons such as hexanes.](#) ~~Because organolithiums are extremely air- and moisture-sensitive, proper care must be taken to handle these reagents.~~ *N*-BuLi can be used as a base or nucleophile depending on the reaction conditions. [Furthermore, it can be used in halogen-lithium exchange reactions to generate aryl or vinyl lithium reagents, which can immediately react with an electrophile. In addition, \*n\*-BuLi is commonly used as an initiator in the polymerization of dienes used in elastomers.](#)

Titration is a common quantitative analytical technique used to determine an unknown concentration of an analyte by adding a known concentration of a titrant until the reaction reaches completion which is indicated by a color change. In this experiment, we will demonstrate the titration of *n*-BuLi and calculate its concentration using diphenylacetic acid as the titrant, which is a stable solid and can be easily weighed. Upon reaction of the *n*-BuLi with diphenylacetic acid, the solution turns a deep yellow color at the end point of the titration, meaning an indicator is not needed (**Figure 1**). By measuring the precise volume of *n*-BuLi used to consume the diphenylacetic acid, we can calculate the molarity of the reaction. This experiment should be repeated three times to obtain an average measurement.

## Procedure:

1. Preparation of Titrant
  - 1.1. To a flame-dried round-bottom flask equipped with a stir bar under N<sub>2</sub> added diphenylacetic acid (250 mg, 1.18 mmol) and anhydrous THF (5 mL).
2. Titration of *n*-BuLi
  - 2.1. Calculate the approximate amount of *n*-BuLi solution (in hexanes) needed to consume the diphenylacetic acid. Using a syringe, slowly add a solution of *n*-BuLi dropwise. The reaction mixture will turn temporarily yellow and back to a colorless solution with each drop.
  - 2.2. Continue adding the solution of *n*-BuLi dropwise until the reaction mixture continues to stay a deep yellow color. This will be the end point of the reaction.
  - 2.3. Repeat the titration 3x to obtain an average volume of *n*-BuLi used.
3. Calculation of Molarity
  - 3.1. The moles of *n*-BuLi used in the titration are equivalent to the moles of diphenylacetic acid used in the reaction. Since the moles of diphenylacetic acid used is known, the calculation of concentration of *n*-BuLi is given by the following equation: mmol diphenylacetic acid/mL *n*-BuLi = molarity of *n*-BuLi solution

## 4. Addition of *n*-BuLi to Benzaldehyde (Figure 2)

- 4.1. To a flame-dried round bottom flask equipped with a stir bar under N<sub>2</sub> add anhydrous THF (30 mL) and benzaldehyde (3 mmol, 1 equiv). Cool the solution to -78 °C.
- 4.2. Add *n*-BuLi (1.1 equiv) and let the reaction warm to room temperature.
- 4.3. Add saturated NH<sub>4</sub>Cl (aq, 10 mL) to quench the reaction and extract the aqueous layer with diethyl ether (2 x 25 mL).
- 4.4. Wash the combined organic layers with water (2 x 15 mL) and brine (1 x 20 mL).
- 4.5. The combined organic layers were dried with Na<sub>2</sub>SO<sub>4</sub>, filtered, and concentrated under reduced pressure to afford the product.
- 3.1.

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## Representative Results:

Representative results for the titration of *n*-BuLi for Procedures 1-3

Procedure Step	Color of reaction mixture
1.1	Colorless
2.1	Yellow, then colorless
2.2	Deep yellow
3.1	1.18 mmol diphenylacetic acid/1 mL <i>n</i> -BuLi = 1.18 M <i>n</i> -BuLi in hexanes

Table 1. Representative results for procedures 1-3.

## Summary:

In this experiment, we have demonstrated how to determine the concentration of an *n*-BuLi solution by using diphenylacetic acid. [We have also performed a reaction by adding \*n\*-BuLi to benzaldehyde.](#)

## Applications

Accurate concentrations of *n*-BuLi are important in order to be used successfully in a number of reactions. *n*-BuLi is commonly used in lithium-halogen exchange reactions to prepare aryllithium or vinylolithium reagents, which can be subsequently used in C-C bond forming reactions. [Furthermore, it can be used as a base or nucleophile, depending on the reaction conditions.](#)

## Legend:

**Figure 1.** Reaction of *n*-BuLi with diphenylacetic acid.

[Figure 2.](#) Addition of *n*-BuLi to benzaldehyde.

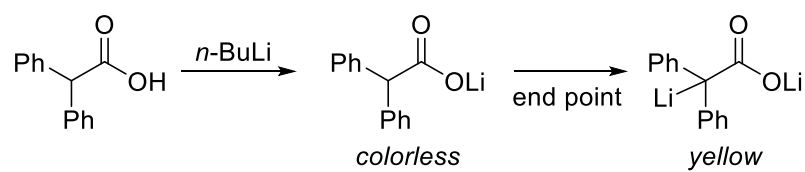


Figure 1

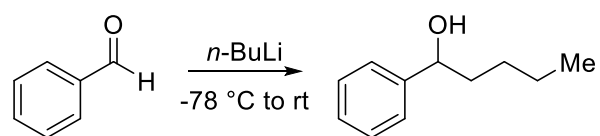


Figure 2